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amounts of substances in a mixture is with the mole fraction
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mixture to the total number of moles of all substances for a
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each would be written as

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alcohol n_b 23 g 46 g mol 1 0 5 mol number of moles of
solvent water n_a 54 g 18 g mol 1 3 mol total number of
moles $n_a + n_b$ 0 5 3 3 5 mol mole fraction of solute ethyl
alcohol $x_b = \frac{n_b}{n_a + n_b} = \frac{0.5}{0.5 + 3.3} = 0.1429$

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and water in a solution prepared by dissolving 4.5 g of
alcohol in 40 g of H_2O molar mass of H_2O is 18 gm mole and
molar mass of CH_3OH is 32 gm mole solution moles of CH_3OH
 $\frac{4.5}{32} = 0.14$ mole moles of H_2O $\frac{40}{18} = 2.2$ moles therefore
according to the equation

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mixtures where the partial pressure of a gas was its mole
fraction time the total pressure $p_a = x_a p_t$ mole fraction a left
frac mole a moles total right sum n_1 all $x_i = \frac{n_i}{n_1}$ mole fraction
of substance a is often denoted by x_a and

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web notice that the mole fraction has no units on it and is written as a decimal value do not change it to percent note of caution you could see the term mole percent it is simply the mole fraction multiplied by 100 for example in the problem just below the mole fraction of cinnamic acid is 0.2885 its mole percent would be 28.85

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web since the mole fraction of the o₂ is 0.050 it contributes 0.050 or if you will 5.0% of the total pressure thus total pressure is 0.210 0.050 4.2 atm problem 9 a sample of 1.43 g of helium and an unweighed quantity of o₂ are mixed in a flask at room temperature